

NO:	FINDING EMPIRICAL FORMULA-SET-1															
1	<p>Analysis of a compound formed between magnesium and nitrogen showed it contained 14.4 g of magnesium and 5.6 g of nitrogen.</p> <p>What is the empirical formula of the compound?</p> <p><b>A</b> Mg<sub>2</sub>N<sub>3</sub>      <b>B</b> Mg<sub>3</sub>N<sub>2</sub>      <b>C</b> Mg<sub>4</sub>N<sub>6</sub>      <b>D</b> Mg<sub>6</sub>N<sub>4</sub></p>															
Ms-1	<b>B</b>															
2	<p>A compound is analysed and found to contain 85.7% carbon and 14.3% hydrogen.</p> <p>What is its empirical formula?</p> <p><b>A</b> CH      <b>B</b> CH<sub>2</sub>      <b>C</b> C<sub>2</sub>H<sub>4</sub>      <b>D</b> C<sub>6</sub>H</p>															
Ms-2	<b>B</b>															
3	<p>The relative molecular mass of an alcohol is 88.</p> <p>Its percentage composition by mass is: C, 54.5%; H, 9.1%; O, 36.4%.</p> <p>Which row shows the empirical formula and molecular formula for this alcohol?</p> <table border="1" data-bbox="272 1209 889 1472"> <thead> <tr> <th></th> <th>empirical formula</th> <th>molecular formula</th> </tr> </thead> <tbody> <tr> <td><b>A</b></td> <td>C<sub>2</sub>H<sub>4</sub>O</td> <td>C<sub>2</sub>H<sub>4</sub>O</td> </tr> <tr> <td><b>B</b></td> <td>C<sub>2</sub>H<sub>4</sub>O</td> <td>C<sub>4</sub>H<sub>8</sub>O<sub>2</sub></td> </tr> <tr> <td><b>C</b></td> <td>C<sub>4</sub>H<sub>8</sub>O<sub>2</sub></td> <td>C<sub>4</sub>H<sub>8</sub>O<sub>2</sub></td> </tr> <tr> <td><b>D</b></td> <td>C<sub>4</sub>H<sub>8</sub>O<sub>2</sub></td> <td>C<sub>2</sub>H<sub>4</sub>O</td> </tr> </tbody> </table>		empirical formula	molecular formula	<b>A</b>	C <sub>2</sub> H <sub>4</sub> O	C <sub>2</sub> H <sub>4</sub> O	<b>B</b>	C <sub>2</sub> H <sub>4</sub> O	C <sub>4</sub> H <sub>8</sub> O <sub>2</sub>	<b>C</b>	C <sub>4</sub> H <sub>8</sub> O <sub>2</sub>	C <sub>4</sub> H <sub>8</sub> O <sub>2</sub>	<b>D</b>	C <sub>4</sub> H <sub>8</sub> O <sub>2</sub>	C <sub>2</sub> H <sub>4</sub> O
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Ms-3	<b>B</b>															
4	<p>A compound contains 34.5% calcium, 24.1% silicon and 41.4% oxygen by mass.</p> <p>What is its empirical formula?</p> <p><b>A</b> Ca<sub>2</sub>SiO<sub>3</sub>      <b>B</b> CaSiO<sub>3</sub>      <b>C</b> CaSi<sub>2</sub>O<sub>3</sub>      <b>D</b> CaSiO<sub>6</sub></p>															
Ms-4	<b>B</b>															